**1.** How many bridging oxygen atoms are present in $P_4O_{10}$?

(1)6 (2)4 (3) 2 (4) 5

**2.** Which of the following phosphorus is the most reactive?

(1) Scarlet phosphorus (2) White phosphorus

(3) Red phosphorus (4) Violet phosphorus

**3.** Which of the following can be hydrolysed?

(1)$TeF_6$ (2) $NCl_3$

(3) $SF_6$ (4) Both 1 & 2

**4.** Which of the following methyl diboranes does not exist ?

(1) $B_2H_4(CH_3)_2$ (2) $B_2H_5(CH_3)$

(3) $B_2H_2(CH_3)_4$ (4) $B_2H(CH_3)_5$

**5.** Which of the following is a sesquioxide?

(1) $N_2O_4$ (2) $N_2O_3$

(3) $N_2O$ (4) $N_2O_5$

**6.** Which of the following does not form during hydrolysis of $XeF_6$?

(1) $XeO_3$ (2) $XeOF_4$

(3) $XeO_2F_2$ (4) $XeOF_3$

**7.** Mixture used in Holme's signal is?

(1) $CaC_2$ and $CaCl_2$ (2) $CaCl_2$ and $Ca_3P_2$

(3) $CaC_2$ and $Ca_3N_2$ (4) $CaC_2$ and $Ca_3P_2$

**8.** $HNO_2$ acts as an/a

(1) acid (2) oxidising agent

(3) reducing agent (4) All

**9.** The nitrogen oxide that does not contain N-N bond are-

(1) $N_2O$ (2) $N_2O_3$

(3) $N_2O_4$ (4) $N_2O_5$

**10.** Bleaching action of $SO_2$ is due to–

(1) reduction (2) oxidation

(3) hydrolysis (4) acidic nature

**11.** $PH_3$ can be formed in

(1) hydrolysis of calcium phosphide

(2) heating $H_3PO_3$

(3) reaction of $P_4$ & NaOH in inert atmosphere

(4) All

**12.** A (black compound) $\xrightarrow{\text{halogen acid}}$ B

(green yellow gas)

B(excess) $\xrightarrow{NH_3}$ C (unstable trihalide)

Correct statement is

(1) C is $PCl_3$ (2) halogen acid is HI

(3) B is $Cl_2$ (4) A is $KMnO_4$

**13.** Beryl is a type of

(1) chain silicate (2) cyclic silicate

(3) sheet silicate (4) 3-D silicate

**14.** Order of $pK_a$

(1) $HOCl > HClO_2 > HClO_3 > HClO_4$

(2) $HClO_4 > HClO_3 > HClO_2 > HOCl$

(3) $HOCl > HClO_3 > HClO_4 > HOCl$

(4) $HClO_3 > HClO_2 > HOCl > HClO_4$

**15.** Which of the following compound will not give $NH_3$ on heating

(1) $(NH_4)_2SO_4$ (2) $(NH_4)_2CO_3$

(3) $NH_4NO_2$ (4) $NH_4Cl$

**16.** On hydrolysis $CaC_2$ gives a gas which on trimerisation gives

(1) $C_2H_2$ (2) $C_6H_6$

(3) $C_2H_4$ (4) $C_3H_8$

**17.** $P+Cl_2 \rightarrow A$, $P$ + Excess $Cl_2 \rightarrow B$

Hydrolysis products of A and B are respectively

(1) $H_3PO_2$, $H_3PO_3$ (2) $H_3PO_4$, $H_3PO_3$

(3) $H_3PO_3$, $H_3PO_4$ (4) $H_3PO_2$, $H_3PO_4$

**18.** Which among the following order of given properties is incorrect

(1) $HOCl > HClO_2 > HClO_3 > HClO_4$ – oxidising nature

(2) $Cl > F > Br > I$ – electron affinity

(3) $Cl_2 > Br_2 > I_2 > F_2$ – bond dissociation energy

(4) $HF < HCl < HBr < HI$ – acidic nature

**19.** Amongst the following the central atoms are directly bonded in

(1) $N_2O_5$ (2) $S_2O_5^{2-}$

(3) $P_4O_{10}$ (4) $Mn_2O_7$

**20.** The chain length of silicones can be controlled by

(1) $(CH_3)_3SiCl$

(2) Addition of Cu powder

(3) Elevation of tempeature

(4) None of these

**21.** Which of the following statement is incorrect regarding $B_2H_6$?

(1) On methylation it gives $B_2H(CH_3)_5$

(2) It has two $2e^-3C$ bonds

(3) There can be maximum six atoms in a plane

(4) It has four $2e^-2C$ bonds

**22.** Red and white phosphorus are similar in

(1) smell (2) solubility in $CS_2$

(3) Hybridisation of P (4) Stability

**23.** Which of the following is strongest oxidizing agent

(1) HOCl (2) $HClO_2$

(3) $HClO_3$ (4) $HClO_4$

**24.** In which species O–O bond is present

(1) $S_2O_8^{-2}$ (2) $S_4O_6^{-2}$

(3) $SO_3^{-2}$ (4) $S_2O_7^{-2}$

**25.** Glass is soluble in

(1) aqua ragia (2) $H_2SO_4$

(3) HF (4) $HClO_4$

**26.** Paramagnetic oxide is

(1) NO (2) $N_2O_4$

(3) $P_4O_6$ (4) $N_2O_5$

**27.** Which oxide does not act as a reducing agent

(1) NO (2) $NO_2$

(3) $N_2O$ (4) $N_2O_5$

**28.** Borax bead test is not given by

(1) An aluminium salt (2) A cobalt salt

(3) A copper salt (4) A nickel salt

**29.** The reaction showing endothermic nature and reduction of halogen is

(1) $F_2 + ½O_2 \rightarrow F_2O$

(2) $Cl_2 + O_2 \rightarrow Cl_2O$

(3) $F_2 + H_2O \rightarrow 2HF + ½O_2$

(4) None of the above

**30.** In which of the following oxy acid of sulphur sulphur atoms has different oxidation states

(1) $H_2S_4O_6$ (2) $H_2S_2O_8$

(3) $H_2S_2O_4$ (4) All

**31.** Which of the following is most acidic?

(1) $Cl_2O_7$ (2) $SO_3$

(3) $P_2O_5$ (4) $SiO_2$

**32.** Which of the following is not peroxide

(1) $Na_2O_2$ (2) $CaO_2$

(3) $PbO_2$ (4) $H_2O_2$

**33.** In which of the following reaction, phosphine is not obtained as product

(1) $Ca_3P_2 + HCl \longrightarrow$ (2) $P_4 + NaOH \longrightarrow$

(3) $H_3PO_4 \xrightarrow{\Delta}$ (4) $H_3PO_3 \xrightarrow{\Delta}$

**34.** Correct statement about boric acid is

(1) boron is $sp^3$ hybridised

(2) boric acid is triprotic acid

(3) it is used in the treatement of eye infection

(4) It forms covalent network

**35.** Which of the following oxide is acidic in nature?

(1) $B_2O_3$ (2) $Al_2O_3$ (3) $Ga_2O_3$ (4) $In_2O_3$

**36.** The most commonly used reducing agent is

(1) $AlCl_3$ (2) $PbCl_2$ (3) $SnCl_4$ (4) $SnCl_2$

**37.** Which of the following are peroxy-acid of sulphur

(1) $H_2SO_5$ and $H_2S_2O_8$

(2) $H_2SO_5$ and $H_2S_2O_7$

(3) $H_2S_2O_7$ and $H_2S_2O_8$

(4) $H_2S_2O_6$ and $H_2S_2O_7$

**38.** The element which forms neutral as well as acidic oxide is :-

(1) Sn (2) Si (3) C (4) P

**39.** $P_4O_{10}$ has short and long P–O bonds. The number of short P–O bonds in this compound is :-

(1) 1 (2) 2 (3) 3 (4) 4

**40.** Nitrogen dioxide can be obtained by heating:-

(1) $KNO_3$ (2) $Pb(NO_3)_2$

(3) $NH_4NO_3$ (4) All of these

**41.** $MF + XeF_4 \rightarrow$ 'A' (M – alkali metal) hybridisation of A and shape of A is :-
(1) $sp^3d$, Trigonal bipuramidal
(2) $sp^3d^3$, distoted octahedral
(3) $sp^3d^3$, pentagonal planer
(4) No compound formed at all

**42.** Which oxide is most acidic :-
(1) $Al_2O_3$ (2) $Na_2O$
(3) MgO (4) CaO

**43.** $Cl_2O_6$ is an anhydride of :-
(1) $HClO_3$
(2) $HClO_2$
(3) $HClO_4$
(4) Mixed anhydride of $HClO_3$ & $HClO_4$

**44.** Carbogen is a mixture of
(1) $O_2$ & $H_2$ (2) $CO_2$ & $O_2$
(3) $O_2$ & Air (4) $O_2$ & Ne

**45.** A black sulphide when treated with ozone becomes white, the white compound is :-
(1) $ZnSO_4$ (2) $CaSO_4$
(3) $BaSO_4$ (4) $PbSO_4$

**46.** An example of tetrabasic acid is :-
(1) Orthophosphorus acid
(2) Orthophophoric acid
(3) Metaphosphoric acid
(4) Pyrophosphoric acid

**47.** The reducing power of divalent species decreases in the order is :-
(1) Ge > Sn > Pb (2) Sn > Ge > Pb
(3) Pb > Sn > Ge (4) None of these

**48.** Basicity of phosphinic acid is
(1) 1 (2) 2 (3) 3 (4) 4

**49.** Empirical formula of Bleaching powder is :-
(1) $Ca(OH)_2$ (2) $CaOCl_2$
(3) $Ca(OCl)_2$ (4) $CaClO_3$

**50.** Which of the following silicate is called disilicate?
(1) Orthosilicate (2) Pyrosilicate
(3) Single-chain silicate (4) None of these

**51.** The cationic part of solid $XeF_6$ is having the "_____" shape:
(1) Linear (2) Angular
(3) Square pyramidal (4) Tetrahedral

**52.** In the hydrolysis of ICl, the products are:
(1) HI + HCI (2) HI + HOCI
(3) HCI + HOI (4) HOCI + HOI

**53.** Which of the following compounds are the common products obtained in the hydrolysis of $XeF_6$ and $XeF_4$?
(1) $XeO_2F_2$ (2) HF
(3) $XeO_3$ (4) Both (2) and (3)

**54.** Find the incorrect match:
(1) $Al_2Cl_6$ : 3C-4e bond is present
(2) $Al_2(CH_3)_6$ : All carbon atoms are $sp^3$-hybridized
(3) $I_2Cl_6$ : Nonplanar
(4) $Al_2Br_6$ : Nonpolar

**55.** Which of the following is not a Lewis acid?
(1) $SiF_4$ (2) $FeCl_3$ (3) $BF_3$ (4) $C_2H_4$

**56.** Thallium shows different oxidation states because:
(1) Of its high reactivity
(2) Of inert pair of electrons
(3) Of its amphoteric nature
(4) It is a transition metal

**57.** $H_3BO_3$ is :
(1) Monobasic and weak Lewis acid
(2) Monobasic and weak bronsted acid
(3) Monobasic acid and strong Lewis acid
(4) Tribasic acid and weak bronsted acid

**58.** What is formula for carbon suboxide?
(1) CO (2) $CO_2$
(3) $C_2O_4$ (4) $C_3O_2$

**59.** $CCl_4$ is used as fire extinguisher because:
(1) Its m.pt. is high
(2) It forms covalent bond
(3) Its b.pt. is low
(4) It gives incombustible vapours

**60.** Marsh gas contains:
(1) $CH_4$ (2) $CO_2$
(3) $C_2H_6$ (4) $N_2$

**61.** Carborundum is:
(1) $Al_2O_3$ (2) SiC (3) $BF_3$ (4) $B_4C$

**62.** The acid which contains a peroxo linkage is:
(1) Sulphurous acid (2) Pyrosulphuric acid
(3) Dithionic acid (4) Caro's acid

**63.** In $SiF_6^{2-}$ and $SiCl_6^{2-}$, which one is known and why?

(1) $SiF_6^{2-}$ because of small size of F

(2) $SiF_6^{2-}$ because of large size of F

(3) $SiCl_6^{2-}$ because of small size of Cl

(4) $SiCl_6^{2-}$ because of large size of Cl

**64.** The stability of dihalides of Si, Ge, Sn, and Pb increase stability in the sequence:

(1) $PbX_2 < SnX_2 < GeX_2 < SiX_2$

(2) $GeX_2 < SiX_2 < SnX_2 < PbX_2$

(3) $SiX_2 < GeX_2 < PbX_2 < SnX_2$

(4) $SiX_2 < GeX_2 < SnX_2 < PbX_2$

**65.** Products formed when $Pb(NO_3)_2$ is heated are:

(1) PbO, $N_2$,$O_2$ (2) $Pb(NO_2)_2$,$O_2$

(3) PbO,$NO_2$,$O_2$ (4) Pb, $N_2$, $O_2$

**66.** Silver chloride dissolves in excess of $NH_4OH$. The cation present in solution is:

(1) $Ag^+$ (2) $[Ag(NH_3)_4]^+$

(3) $[Ag(NH_3)_2]^+$ (4) $[Ag(NH_3)_6]^+$

**67.** The catalyst used in the manufacture of ammonia by Haber's process is:

(1) Pt (2) $Fe_2O_3$ (3) $CuCl_2$ (4) $V_2O_5$

**68.** Industrial preparation of nitric acid by Ostwald's process involves:

(1) Oxidation of $NH_3$

(2) Reduction of $NH_3$

(3) Hydrogenation of $NH_3$

(4) Hydrolysis of $NH_3$

**69.** When zinc reacts with very dilute nitric acid it produces:

(1) $NH_4NO_3$ (2) NO

(3) $NO_2$ (4) $H_2$

**70.** Which of the following oxides of nitrogen is the an-hydride of nitrous acid?

(1) NO (2) $N_2O_3$

(3) $N_2O_4$ (4) $N_2O_5$

**71.** Polar oxide of carbon is :-

(1) CO (2) $CO_2$

(3) $C_3O_2$ (4) Both 1 and 3

**72.** Aqueous solution of ammonia consists of :-

(1) $H^+$ only (2) $OH^-$ only

(3) $NH_4^+$ only (4) $NH_4^+$ and $OH^-$

**73.** Phosphine, acetylene and ammonia can be formed by treating water with:

(1) $Mg_3P_2$, $Al_4C_3$, $Li_3N$

(2) $Ca_3P_2$, $CaC_2$, $Mg_3N_2$

(3) $Ca_3P_2$, $Be_2C$, $NH_4NO_3$

(4) $Ca_3P_2$,$Mg_2C_3$,$NH_4NO_3$

**74.** Atoms in $P_4$ molecule of white phosphorus are arranged regularly in the following way:

(1) At the corners of a cube

(2) At the corners of an octahedron

(3) At the corners of a tetrahedron

(4) At the center and corners of a tetrahedron

**75.** Which is formed when $K_2Cr_2O_7$, $CaCl_2$, and conc. $H_2SO_4$ are heated?

(1) $Cr_2(SO_4)_3$ (2) $CrCl_3$

(3) $CrO_2Cl_2$ (4) $K_2CrO_4$

**76.** Which one of the following reactions does not occur ?

(1) $F_2 + 2Cl^- \longrightarrow 2F^- + Cl_2$

(2) $Cl_2 + 2F^- \longrightarrow 2Cl^- + F_2$

(3) $Br_2 + 2I^- \longrightarrow 2Br^- + I_2$

(4) $Cl_2 + 2Br^- \longrightarrow 2Cl^- + Br_2$

**77.** $BCl_3$ does not exist as dimer but $BH_3$ exists as dimer ($B_2H_6$) because:

(1) Chlorine is more electronegative than hydrogen

(2) There is $p_\pi - p_\pi$ back bonding in $BCl_3$, but $BH_3$ does not contain such multiple bonding

(3) Large sized chlorine atoms do not fit in between the small boron atoms, whereas small-sized hydrogen atoms fit between boron atoms

(4) None of these

**78.** Red lead is :
(1) PbO (2) $Pb_3O_4$
(3) $PbO_2$ (4) HgS

**79.** The most stable and basic hydride of 15th group is :
(1) $NH_3$ (2) $PH_3$ (3) $AsH_3$ (4) $BiH_3$

**80.** C — C bond length is maximum in:
(1) Diamond (2) Graphite
(3) Napthalene (4) Fullerene

**81.** N forms $NCl_3$, whereas P can form both $PCl_3$ and $PCl_5$. Why?
(1) P has vacant d-orbitals which can be used for bonding but N does not have
(2) N atom is larger in size than P
(3) P is more reactive towards Cl than N
(4) None of the above

**82.** Which is in the decreasing order of boiling points of VA group hydrides?
(1) $NH_3 > PH_3 > AsH_3 > SbH_3$
(2) $SbH_3 > AsH_3 > PH_3 > NH_3$
(3) $PH_3 > NH_3 > AsH_3 > SbH_3$
(4) $SbH_3 > NH_3 > AsH_3 > PH_3$

**83.** Which compound acts as an oxidizing as well as a reducing agent?
(1) $SO_2$ (2) $Mn_2O_7$
(3) $Al_2O_3$ (4) $CrO_3$

**84.** The most powerful oxidizing agent is :
(1) Fluorine (2) Chlorine
(3) Bromine (4) Iodine

**85.** Which one of the hydracids does not form any precipitate with $AgNO_3$?
(1) HF (2) HCl (3) HBr (4) HI

**86.** Which of the following structure is non-planar?
(1) $Na_3B_3O_6$
(2) $I_2Cl_6$
(3) Sheet silicates
(4) Inorganic graphite layer

**87.** In the reaction $LiH + AlH_3 \longrightarrow LiAlH_4$ $AlH_3$ and LiH act as:
(1) Lewis acid and Lewis base
(2) Lewis base and Lewis acid
(3) Bronsted base and Bronsted acid
(4) None of these

**88.** Which one is not an acid salt?
(1) $NaH_2PO_2$ (2) $NaH_2PO_3$
(3) $NaH_2PO_4$ (4) None of these

**89.** The most thermodynamically stable allotropic form of phosphorus is:
(1) Red (2) White
(3) Black (4) Yellow

**90.** Identify the incorrect statement among the following:
(1) Ozone reacts with $SO_2$ to give $SO_3$
(2) Silicon reacts with NaOH(aq.) in the presence of air to give $Na_2SiO_3$
(3) $Cl_2$ reacts with excess of $NH_3$ to give $N_2$ and $NH_4Cl$.
(4) $Br_2$ reacts with hot and strong NaOH solution to give NaBr, $NaBrO_4$ and $H_2O$.

**91.** The product of oxidation of $I^-$ with $MnO_4^-$ in alkaline medium is:
(1) $IO_3^-$ (2) $I_2$ (3) $IO^-$ (4) $IO_4^-$

**92.** The chemical formula of feldspar is :
(1) $KAlSi_3O_8$ (2) $Na_3AlF_6$
(3) $NaAlO_2$ (4) $K_2SO_4$

**93.** Which of the following is correct match:
(1) Gun metal → Cu, Sn & Zn (Red brass)
(2) White metal → Contains Li
(3) Stainless Steel → Fe, Cr, Ni, Carbon
(4) All

**94.** Which is used for estimation of carbonmonooxide ?
(1) $I_2O_5$ (2) $ClO_2$ (3) $BrO_3$ (4) $Cl_3O_7$

**95.** There is no S—S bond in
(1) $S_2O_4^{2-}$ (2) $S_2O_5^{2-}$
(3) $S_2O_3^{2-}$ (4) $S_2O_7^{2-}$

**96.** Total number of lone pairs of electron and P–O–P linkage present in dimer of $P_2O_5$ are :
(1) 16 (2) 22 (3) 26 (4) 30

**97.** The number of S—S bonds in polythionic acid ($H_2S_nO_6$):
(1) n (2) n – 1
(3) n – 2 (4) None of these

**98.** Which of the following halogen oxides is ionic?
(1) $I_4O_9$ (2) $I_2O_5$ (3) $BrO_2$ (4) $C\ell O_3$

**99.** Antichlor is a compound:
(1) Which absorbs chlorine
(2) Which removes excess of $Cl_2$ from a material
(3) Which liberates $Cl_2$ from bleaching powder
(4) Which acts as a catalyst in the manufacture of $Cl_2$

**100.** Which of the following statements regarding orthoboric acid ($H_3BO_3$) is false?
(1) It acts as a weak monobasic acid
(2) It is soluble in hot water
(3) It has a planar structure
(4) It acts as a tribasic acid

**101.** Which of the following oxides is acidic in nature?
(1) $B_2O_3$ (2) $Al_2O_3$ (3) $Ga_2O_3$ (4) ZnO

**102.** Catenation i.e., linking of similar atoms depends on size and electronic configuration of atoms. The tendency of catenation in group 14 elements follows the order
(1) C > Si > Ge > Sn (2) C > > Si > Ge ≈ Sn
(3) Si > C > Sn > Ge (4) Ge > Sn > Si > C

**103.** Cement, the important building material is a mixture of oxides of several elements. Besides calcium, iron and sulphur, oxides of elements of which of the group(s) are present in the mixture?
(1) Group 2
(2) Groups 2, 13 and 14
(3) Groups 2 and 13
(4) Groups 2 and 14

**104.** The possible oxidation state of Tl are:
(1) + 1 and +2 (2) +2 and +3
(3) +1 and – 1 (4) + 1 and + 3

**105.** Nitrogen gas is liberated by thermal decomposition of :
(1) $NH_4NO_2$ (2) $NaN_3$
(3) $(NH_4)_2Cr_2O_7$ (4) All

**106.** $BX_3 + NH_3 \xrightarrow{R.T.} BX_3 \cdot NH_3 + \text{Heat}$ of adduct formation ($\Delta H$), $\Delta H$ is maximum for (Numerical value)
(1) $BF_3$ (2) $BCl_3$
(3) $BBr_3$ (4) $BI_3$

**107.** Which of the following oxyacid contains both P–H and P–P bond simultaneously?
(1) $H_4P_2O_5$ (2) $H_4P_2O_7$
(3) $H_4P_2O_6$ (4) None

**108.** $Na_2B_4O_7 \cdot 10H_2O \xrightarrow{Heat} X + NaBO_2 + H_2O$,
$X + Cr_2O_3 \xrightarrow{Heat} \underset{\text{(Green coloured)}}{Y}$ X and Y are :
(1) $Na_3BO_3$ and $Cr(BO_2)_3$
(2) $Na_2B_4O_7$ and $Cr(BO_2)_3$
(3) $B_2O_3$ and $Cr(BO_2)_3$
(4) $B_2O_3$ and $CrBO_3$



# P-BLOCK ELEMENTS ANSWER KEY

| Que. | 1 | 2 | 3 | 4 | 5 | 6 | 7 | 8 | 9 | 10 | 11 | 12 | 13 | 14 | 15 |
|---|---|---|---|---|---|---|---|---|---|---|---|---|---|---|---|
| Ans. | **1** | **2** | **4** | **4** | **2** | **4** | **4** | **4** | **4** | **1** | **4** | **3** | **2** | **1** | **3** |
| Que. | 16 | 17 | 18 | 19 | 20 | 21 | 22 | 23 | 24 | 25 | 26 | 27 | 28 | 29 | 30 |
| Ans. | **2** | **3** | **3** | **2** | **1** | **1** | **3** | **1** | **1** | **3** | **1** | **4** | **1** | **3** | **1** |
| Que. | 31 | 32 | 33 | 34 | 35 | 36 | 37 | 38 | 39 | 40 | 41 | 42 | 43 | 44 | 45 |
| Ans. | **1** | **3** | **3** | **3** | **1** | **4** | **1** | **3** | **4** | **2** | **3** | **1** | **4** | **2** | **4** |
| Que. | 46 | 47 | 48 | 49 | 50 | 51 | 52 | 53 | 54 | 55 | 56 | 57 | 58 | 59 | 60 |
| Ans. | **4** | **1** | **1** | **2** | **2** | **3** | **3** | **4** | **3** | **4** | **2** | **1** | **4** | **4** | **1** |
| Que. | 61 | 62 | 63 | 64 | 65 | 66 | 67 | 68 | 69 | 70 | 71 | 72 | 73 | 74 | 75 |
| Ans. | **2** | **4** | **1** | **4** | **3** | **3** | **2** | **1** | **1** | **2** | **1** | **4** | **2** | **3** | **3** |
| Que. | 76 | 77 | 78 | 79 | 80 | 81 | 82 | 83 | 84 | 85 | 86 | 87 | 88 | 89 | 90 |
| Ans. | **2** | **3** | **2** | **1** | **1** | **1** | **4** | **1** | **1** | **1** | **3** | **1** | **1** | **3** | **4** |
| Que. | 91 | 92 | 93 | 94 | 95 | 96 | 97 | 98 | 99 | 100 | 101 | 102 | 103 | 104 | 105 |
| Ans. | **1** | **1** | **4** | **1** | **4** | **3** | **2** | **1** | **2** | **4** | **1** | **2** | **2** | **4** | **4** |
| Que. | 106 | 107 | 108 | | | | | | | | | | | | |
| Ans. | **4** | **4** | **3** | | | | | | | | | | | | |